

Chemistry – Chapter 7: Language of Chemistry

Nomenclature Practice Part 2

Name: _____ Date: _____ Period: _____

Covalent Compounds and Acid Nomenclature:

State how a covalent compound can be identified: _____

Explain how naming of a covalent compound differs from that of an ionic compound: _____

State how to identify a simple acid: _____

Name the Covalent Compound From the Given Formula	
Compound	Name
NF₃	Nitrogen trifluoride
SO ₂	
N ₂ O ₄	
SF ₆	
P ₂ O ₅	
NI ₃	
Br ₂ O ₃	
P ₂ F ₄	
SiO ₂	
NI ₃	
P ₄ O ₁₀	
Write the Formula From the Given Name	
Name	Formula
Dinitrogen monoxide	N₂O
Bromine heptafluoride	
Tetraphosphorus octasulfide	
Tetraiodine nonoxide	
Carbon monoxide	
Oxygen difluoride	
Trinitrogen pentoxide	
Pentasulfur hexanitride	

Name the Acid From the Given Formula	
Compound	Name
H_2CrO_4	Chromic acid
H_2SO_3	
HI	
H_2CO_3	
HClO_2	
H_2S	
H_3AsO_4 (AsO_4^{3-} = arsenate)	
$\text{HC}_2\text{H}_3\text{O}_2$	
HBrO_4 (BrO_4^- = perbromate)	
HIO (IO^- = hypoiodite)	
H_3PO_4	
Write the Formula From the Given Name	
Name	Formula
Sulfurous acid	H_2SO_3
Permanganic acid	
Chlorous acid	
Hydrobromic acid	
Nitrous acid	
Sulfuric acid	
Hydrosulfuric acid	
Dichromic acid	

Combination Practice: Give the name or formula as appropriate for the following compounds:

Compound	Type (Ionic, Covalent or Acid)	Answer
XeF_4	Covalent	Xenon tetrafluoride
Hypobromous acid		
PbSe_2		
Tetraarsenic hexoxide		
Stannous acetate		
Br_3O_8		
HCl		
Indium phosphide		