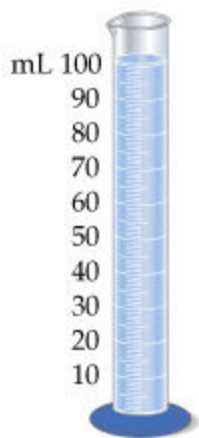
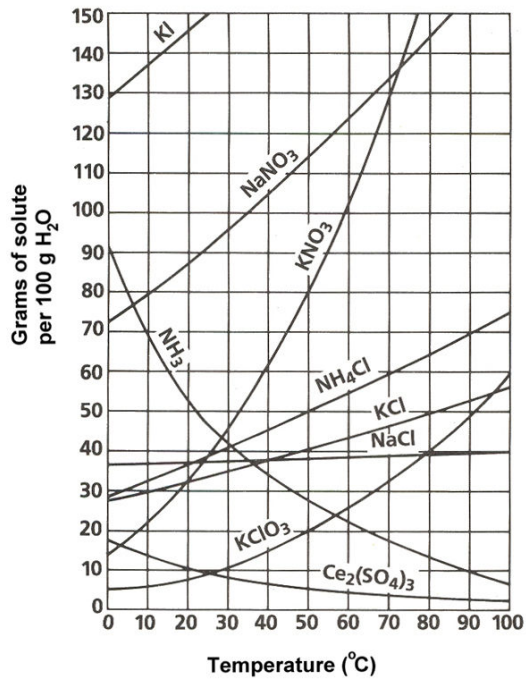


Review: Chapter 1

1. A person runs a fever of 103.5°F . What is this temperature in $^{\circ}\text{C}$?
2. Perform the following calculation to the correct significant digits:
 $(137.5\text{g} + 22.77\text{g} - 3.215\text{g}) / (2.3\text{cm}^3 - 1.77\text{cm}^3) =$
3. Determine the number of inches in 3.75meters.
4. Convert 3.776km to cm and round to 2 significant digits.
5. Classify as a physical or chemical change
 - a. Water is converted to hydrogen and oxygen using electricity.
 - b. Marshmallows are melted to make S'mores.
 - c. Silverware tarnishes.
 - d. An ice skater breaks through the ice.
 - e. Blood is oxygenated
6. Name the following lab item:



7.



Given the solubility curves:

- If 100g of KClO_3 is placed in 300.g of water at $80.^{\circ}\text{C}$, is the solution unsaturated, saturated or oversaturated?
- If the solution is cooled to 30°C , how much, if any KClO_3 will precipitate out of solution (in g)?

8. If a snail moves at 12.3cm/hr , what is this speed in mm/s ?

9. What mass of liquid mercury ($d = 13.6\text{g/cm}^3$) is contained in a volume that is $2.5\text{in} \times 3.5\text{in} \times 1.0\text{in}$?

10. Write the elemental symbols for antimony, cobalt and lead.

Honors Chemistry Chapter 1 Review

1. $^{\circ}\text{C} = 5/9(^{\circ}\text{F}-32)$ $71.5(5/9) = \underline{39.72^{\circ}\text{C}}$

2. $(157.055\text{g}) / (.53\text{cm}^3) = 296.33\text{g}/\text{cm}^3 = \underline{300 \text{ or } 3 \times 10^2 \text{g}/\text{cm}^3}$

3. $3.75\text{m} (100\text{cm}/1\text{m})(1\text{in}/2.54\text{cm}) = 147.638\text{in} = \underline{148\text{in}}$

4. $3.776\text{km} (10^5\text{cm}/1\text{km}) = 377600\text{cm} = \underline{3.8 \times 10^5 \text{cm}}$

5. C, P, C, P, C

6. Graduated cylinder

7. a. Unsaturated. It can hold up to $(40)3 = 120\text{g}$

b. $100\text{g} - 30\text{g} = \underline{70\text{g}}$

8. $12.3\text{cm}/\text{hr} (10\text{mm}/1\text{cm})(1\text{hr}/3600\text{s}) = .034167 = \underline{3.42 \times 10^{-2} \text{mm}/\text{s}}$

9. $(2.5\text{in} \times 3.5\text{in} \times 1.0\text{in})(2.54\text{cm}/1\text{in})^3 (13.6\text{g}/\text{cm}^3) = 1950.06\text{g} = \underline{2.0 \times 10^3 \text{g}}$

10. Sb, Co, Pb