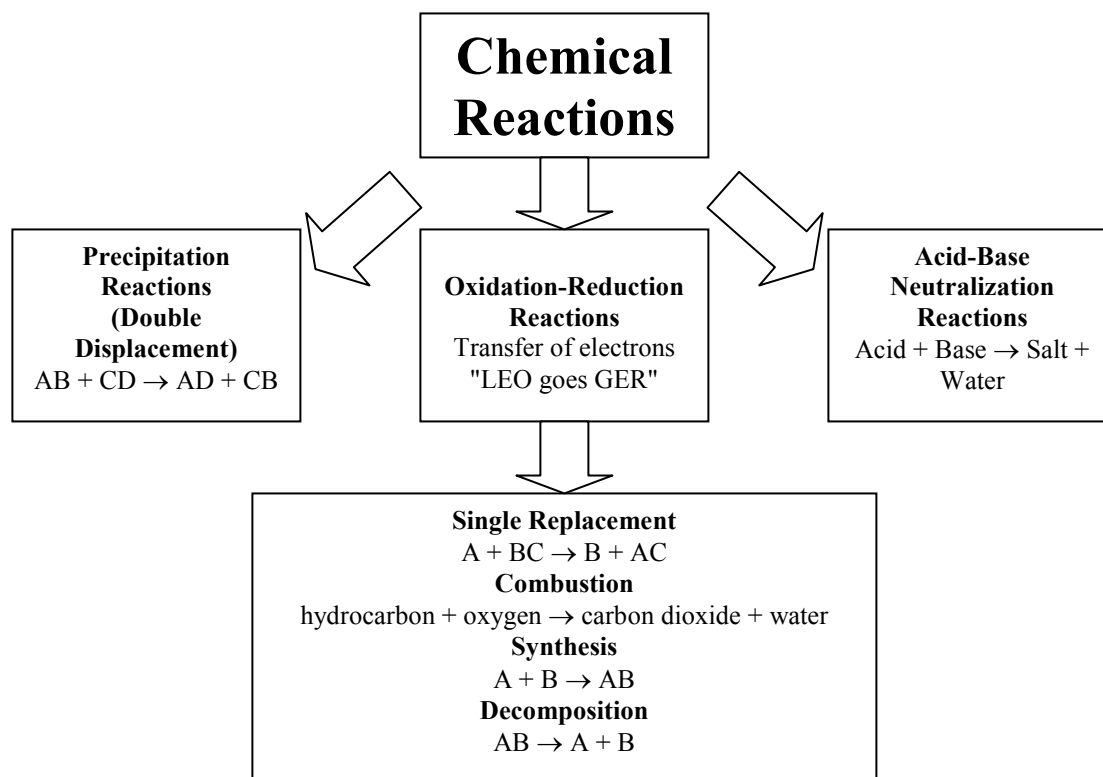


Chemistry

Reaction Types Worksheet chmrxnws.doc

Name _____ Date _____ Period _____

Information:



Questions:

Balance and identify the chemical equations in the table below:

| Balance | Types (List all that apply) | Element oxidized (if redox) | Element reduced (if redox) | Salt formed (if acid-base neutralization) | Precipitate (if double displacement) |
|---|-----------------------------|-----------------------------|----------------------------|---|--------------------------------------|
| $\text{Na} + \text{Br}_2 \rightarrow \text{NaBr}$ | | | | | |
| $\text{C}_3\text{H}_{12} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$ | | | | | |
| $\text{AgNO}_3 + \text{Na}_2\text{S} \rightarrow \text{Ag}_2\text{S} + \text{NaNO}_3$ | | | | | |
| $\text{Al}_2\text{O}_3 \rightarrow \text{Al} + \text{O}_2$ | | | | | |
| $\text{HMnO}_4 + \text{KOH} \rightarrow \text{KMnO}_4 + \text{H}_2\text{O}$ | | | | | |

| | | | | | |
|---|--|--|--|--|--|
| $K + S_8 \rightarrow K_2S$ | | | | | |
| $Ca + O_2 \rightarrow CaO$ | | | | | |
| $C_2H_6O + O_2 \rightarrow CO_2 + H_2O$ | | | | | |
| $(NH_4)_2CO_3 + MgCl_2 \rightarrow NH_4Cl + MgCO_3$ | | | | | |
| $HBr + NaOH \rightarrow NaBr + H_2O$ | | | | | |
| $AlCl_3 \rightarrow Al + Cl_2$ | | | | | |
| $H_2O_2 \rightarrow H_2O + O_2$ (Note: O is a 1 ⁻ in hydrogen peroxide) | | | | | |
| $H_2 + O_2 \rightarrow H_2O$ | | | | | |
| $Al + CuBr_2 \rightarrow Cu + AlBr_3$ | | | | | |
| $AlN \rightarrow Al + N_2$ | | | | | |
| $Cu + AgNO_3 \rightarrow Cu(NO_3)_2 + Ag$ | | | | | |